**I.P.S.Sr.Sec.School**

**Max Time : 1 hr** **Class : 11th Chemistry Max Marks : 30**

**Unit Test**

1. Write IUPAC name of the following: [ 1 x 25 = 25 ]

1. CH3 – C C – CH2 – C CH 2. (CH3)2 CH – CH CH2 3. (CH3)3 C – CH CH2

4.  5. 

6.  7. 

8.  9. 

10.  11. 

12.  13. 

14.  15.  16. 

17.  18.  19. 

20.  21. 

22.  23. 

24.  25. 

1. Calculate the pH value of a 0.01 N solution of acetic acid. Ka for CH3COOH is 1.8 x 10 – 5 at 25˚C. [ 2 ]
2. Calculate the pH value of [ 3 ]

|  |  |  |
| --- | --- | --- |
| (i) 0.0001 M NaOH | (ii) 0.01 M NaOH | (iii) 0.04 M NaOH solution at 25˚C. |